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Question 1 (a) (i) Pb^{2+}
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2COOH(aq) + H 2O(l)

CH 3CH 2COO (aq) +

H 3O+(aq) 2. Propanoic
acid, CH 3CH 2COOH,

is a carboxylic acid that

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reacts with water
according to the
equation above. At

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(d) Calculate the number of moles of precipitate that is produced in the experiment. $1.698 \text{ g} / 1.462 \text{ g} = 0.236 \text{ g PbI}_2$
(s) $2 \text{ mol PbI}_2 / 0.236 \text{ g PbI}_2 = 5.12 \text{ mol PbI}_2$

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current exam, but the question types and the topics are similar,

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(numbers in parentheses
and red, represent
guesstimate points)

Question 1 (a) (i) Pb^{2+}

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(1) (ii) Potassium and nitrate ions are in solution at the beginning of the reaction and at the end, do not change in any way (i.e., they are spectators), and as such can be omitted from the equation.

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AP Chemistry exam will
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For those of you who
took the AP Bio exam in
2013, the format change
is the same; more
theoretical, and less

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emphasis on formulas.

We took the first half of the 2014 AP Chemistry College Board Practice Test and it was a lot different from a released AP Chemistry Test from another year...

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