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~~Determination of
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~~(Week 3 \u0026 4)~~

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~~Finding and~~

~~Calculating an
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~~of a Compound |~~

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Experient 9 - The
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Answers to Post-
Lab Questions for
Experiment 5. I
have 40 L of
methane gas
measured at
standard
temperature and
pressure (STP).

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How many moles of methane gas is this?
1 mole of any gas at STP = 22.4 L

Moles of methane = $(40 / 22.4) = 1.79$ moles. What is the volume of 0.30 moles of nitrogen gas if measured at STP? 1 mole of any gas at STP = 22.4 L
Volume of nitrogen = $0.3 \times$

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22.4 = 6.72 L.

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[Copper metal to
Copper (11)
nitrate) 2pts.
Balanced chemical

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your observations
below after

watching the video
titled Formation of
copper (1) nitrate.

Part 2: [
Preparation of
Copper (11)
hydroxide] 2pts.

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First calculate the
moles of KHP

(potassium
hydrogen phthalate,
 $\text{KHC}_8\text{H}_4\text{O}_4$). The
molar mass for KHP
= 204.2 g/mole.

0.531 g KHP x (1

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mole KHP / 204.2 g
KHP) = 0.00260
moles KHP. KHP...

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precipitate most
likely came from
the presence of
chloride ions. Ag^+
 $+ \text{Cl}^- \rightarrow \text{AgCl}(s)$

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water as products?

Yes: it's a
combustion reaction

2) Does your
reaction have two
(or more)

chemicals
combining to form
one chemical? Yes:
it's a synthesis
reaction

~~Chemistry of
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