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following problems. Round your answers using significant figures. 1) Calculate the mass of 15.0 L of NH_3 at 27°C and 900.0 mm Hg . 2) A volume of 26.5 mL of nitrogen gas was collected in a tube at a temperature of 17°C and a pressure of 737 ...

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Mixed Gas Laws Worksheet Answers

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Solutions 1) How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K? $n = PV = (2.8 \text{ atm})(98 \text{ L}) = 11$

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moles of gas RT (0.0821 L.atm/mol.K)(292 K) 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25 0

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Solutions 1) How many moles of gas

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occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K? $n = \frac{PV}{RT} = \frac{(2.8 \text{ atm})(98 \text{ L})}{(0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})} = 11$ moles of gas

2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 25 °C

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The Mixed Gas Laws Worksheet

Answers will explain the following: how much is in a tank, what type of fuel is used, how and where it is stored, and when it is available to be used. The

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answers that you will receive for these questions will vary from one state to another, but you will most likely receive similar answers.

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a temperature of 292 K? $n = PV = (2.8 \text{ atm})(98 \text{ L}) = 11 \text{ moles of gas}$
 $RT (0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})$ 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are

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$RT (0.0821 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K})(292 \text{ K})$ 2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a

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2) If 5.0 moles of O_2 and 3.0 moles of N_2 are placed in a 30.0 L tank at a temperature of 250

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